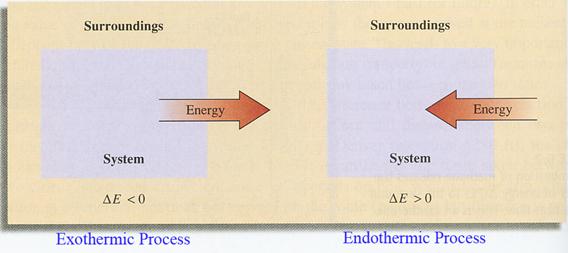
**Unit 7 Notes**

**Endothermic and Exothermic**

In order to consider the energy flow as \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of a chemical reaction we will look at two specific parts: the system and the surroundings

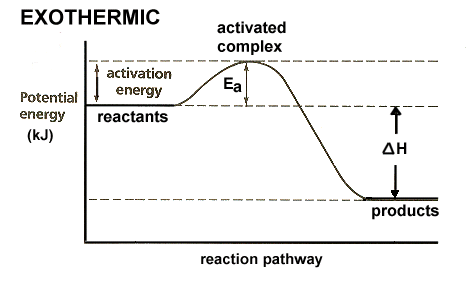
The \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_will be the area of focus, usually the reactants and products of a chemical reaction

The \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ are everything else in the universe other than the system



Exothermic Reactions:

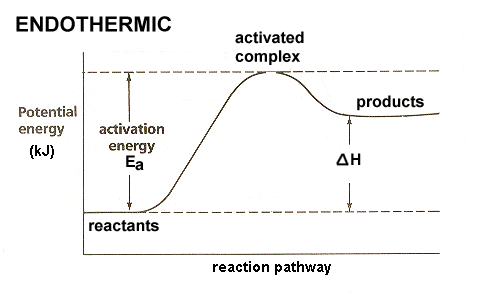
* Energy as heat \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ into the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ by the system.
* Energy written as a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ in a chemical equation. Example: A + BC 🡪 AC + B + Heat
* The system feels \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ to the observer.
* ΔH is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, ΔH is energy that flows as heat



Energy Diagram\*

Endothermic Reactions:

* Energy in the form of heat \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_by the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ from the surroundings
* Energy written as a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ in a chemical equation. Example: AB + CD + heat 🡪 AD + CB
* System feels \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ to the observer
* ΔH is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_



Energy Diagram\*