Station 1: Aluminum Oxide

1. Ionic or covalent
2. Create a neutral compound
3. Draw the Lewis Dot Structure

Station 2:

PI3

1. Ionic or covalent
2. Name the compound
3. Draw the Lewis Dot Structure

Station 3:

Dicarbon Tetrabromide

1. Ionic or covalent
2. Write the formula
3. Draw the Lewis Dot Structure

Station 4:

FePO3

1. Ionic or covalent
2. Determine the charge on the transition metal
3. Write the name of the compound

Station 5:

Dihydrogen monoxide

1. Ionic or Covalent
2. Write the formula
3. Draw the Lewis Dot Structure

Station 6:

Manganese (III) Iodide

1. Ionic or covalent
2. Explain your answer to a (full sentence)
3. Write the formula

Station 7:

a. N3I7 b. C5Se10

c. Si6I9 d. P4S8

Name each of the covalent compounds above.

Station 8:

a. Na2CO3 b. Cr2S3

Name the two ionic compounds above

Write the formulas for the ionic compounds below

c. Potassium Oxide

1. Beryllium Citrate

|  |  |  |  |
| --- | --- | --- | --- |
| 1 | a | b | c |
| 2 | a | b | c |
| 3 | a | b | c |
| 4 | a | b | c |

|  |  |  |  |
| --- | --- | --- | --- |
| 5 | a | b | c |
| 6 | a | b | c |
| 7 | a | b | c | d |
| 8 | a | b | c | d |