**Unit 5 Test Review**

1. Vocabulary to define:
   1. Mole
   2. Avogadro’s number
   3. Molar mass
2. What are the units for molar mass?
3. Calculate the molar mass of the following compounds:
   1. FeSO4
   2. (CH2OH)2
   3. Fe(NO3)3
   4. MgCl2
4. Calculate the percent composition for the following compounds:
   1. MnCl2
   2. Ni(NO3)2
   3. KBrO3
   4. Al2(SO4)3
5. Complete the following conversions:
   1. How many grams are there in 2.3 x 1024 atoms of silver?
   2. How many moles are in 98.3 grams of aluminum hydroxide, Al(OH)3?
   3. How many molecules are in 1.24 mol Cl2?
   4. How many moles are in 1.2 x 103 grams of ammonia, NH3?
   5. How many molecules are there in 450 grams of Na2SO4?
   6. How many grams are there in 7.4 x 1023 molecules of AgNO3?
   7. How many grams are in 0.02 moles of beryllium iodide, BeI2?
   8. How many moles are in 1.20 x 1025 atoms of Phosphorous?
   9. How many moles are in 3.87 x 1021 molecules of AlF?
   10. How many grams are in 4.5 moles of sodium fluoride, NaF?